
CHEMISTRY

9701/33

Paper 3 Advanced Practical Skills 1

October/November 2016

MARK SCHEME

Maximum Mark: 40

Published

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge will not enter into discussions about these mark schemes.

Cambridge is publishing the mark schemes for the October/November 2016 series for most Cambridge IGCSE[®], Cambridge International A and AS Level components and some Cambridge O Level components.

Page 2	Mark Scheme	Syllabus	Paper
	Cambridge International AS/A Level – October/November 2016	9701	33

Question	Answer	Marks
1(a)	<p>Initial and final readings and titre value for rough and initial and final reading for two (or more) accurate titrations</p> <p>Appropriate headings and units and the volume of FA 2 added is recorded for each accurate titration. Headings must match readings Initial/start (burette) and reading/volume Final/end (burette) and reading/volume Titre or volume/vol/FA 2 and used/added (<i>not “difference”, “total”, “V”</i>) Units: /cm³ or (cm³) or in cm³ or cm³ for each volume.</p> <p>All accurate burette readings (initial and final) recorded to nearest 0.05 cm³ <i>Do not award this mark if:</i> <i>50(.00) is used as an initial burette reading;</i> <i>more than one final burette reading is 50.(00);</i> <i>any burette reading is greater than 50.(00)</i></p> <p>Final uncorrected titre is within 0.10 cm³ of any previous uncorrected accurate titre.</p>	<p>1</p> <p>1</p> <p>1</p>
<p>Examiner rounds any accurate burette readings to the nearest 0.05 cm³, checks subtractions and then selects the ‘best’ accurate titres using the hierarchy: identical titres; titres within 0.05 cm³; titres within 0.1 cm³; etc., to calculate mean correct to 0.01 cm³.</p> <p>Examiner compares candidate’s titre value with that of the Supervisor.</p>		
	<p>V, VI and VII Award V, VI and VII for $\delta \leq 0.30 \text{ cm}^3$ Award V and VI for $0.30 < \delta \leq 0.50 \text{ cm}^3$ Award V only for $0.50 < \delta \leq 0.80 \text{ cm}^3$</p>	<p>1</p> <p>1</p> <p>1</p> <p>7</p>

Page 3	Mark Scheme	Syllabus	Paper
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Question	Answer	Marks
1(b)	<p>Calculation of the mean</p> <p>Check mean titre is correctly calculated from clearly selected values (ticks or working)</p> <ul style="list-style-type: none"> Candidate must average two (or more) titres where the total spread is $\leq 0.20 \text{ cm}^3$. Working must be shown or ticks must be put next to the two (or more) accurate readings selected. The mean should normally be quoted to 2 dp rounded to the nearest 0.01. [e.g. 26.667 must be rounded to 26.67.] <p>Two special cases where the mean may not be to 2 dp: allow mean to 3 dp only for 0.025 or 0.075 e.g. 26.325; allow mean to 1 dp if all accurate burette readings were given to 1 dp (ignoring initial given as 0) and the mean is exactly correct. [e.g. 26.0 and 26.2 = 26.1 is correct but 26.0 and 26.1 = 26.1 is incorrect.]</p> <p>Do not award this mark if:</p> <ul style="list-style-type: none"> the rough titre was used to calculate the mean; the candidate carried out only 1 accurate titration; burette readings were incorrectly subtracted to obtain any of the accurate titre values; all burette readings (resulting in titre values used in the calculation of the mean) are integers. <p><i>Note: the candidate's mean will sometimes be marked as correct even if it is different from the mean calculated by the examiner for the purpose of assessing accuracy</i></p>	1
1(c)(i)	I Correctly calculates: $\frac{(b)}{1000} \times 0.0200$	1
1(c)(ii) and 1(c)(iii)	II Correctly uses: (i) $\times 5/2$ and (ii) $/ 0.025$ or (ii) $\times 1000/25$	1
1(c)(iv)	Correctly calculates: (iii) $\times 10$ or (ii) $\times 1000/25 \times 10$	1

Page 4	Mark Scheme	Syllabus	Paper
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Question	Answer	Marks
	3 or 4 significant figures in final answers to all parts (<i>minimum 3 parts attempted</i>)	1 4
	Total:	12

Question	Answer	Marks
2(a)	Examiner to calculate 10% and 20% of Supervisor's time and round this to nearest second. Candidate's time compared with supervisor's time. Award 2 marks if time within 10% of supervisor Award 1 mark if time within 20% of supervisor	2
2(b)(i)	Correctly calculates: $2.61 \times 10^{-5} \times$ reaction time from (a)	1

Page 5	Mark Scheme	Syllabus	Paper
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Question	Answer	Marks
2(b)(ii)	Correctly uses: (i) $\times 0.080$ or (i) $\times 80/1000$ and no additional working	1
2(b)(iii) and 2(b)(iv)	Correctly uses: $2 \times \text{ans (ii)}$ and $(\text{iii})/0.020$ or $(\text{iii}) \times 1000/20$ Time recorded to nearest second in (a) and (c) and 2 – 4 sf in all answers in (b) (minimum 3 parts attempted)	1 1 4
2(c)	Examiner calculates ratio of reaction time (a) /reaction time (b) Award if $1.80 \leq \text{ratio} \leq 2.80$	1 1
2(d)(i)	Time is less/shorter because the amount/volume/concentration of thiosulfate/FA 6 is less (ora) Time is approximately half because (the amount/no. of moles/concentration of) the thiosulfate/FA 6 is half .	1 1

Page 6	Mark Scheme	Syllabus	Paper
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Question	Answer	Marks
2(d)(ii)	(No because) the error is greater in (c) with some explanation e.g. because more readings taken/water added	1
	The measuring cylinder is used more times in (c) or smaller volumes /10 cm ³ instead of 20 cm ³ are measured in (c) or 6 rather than 5 readings taken/more reagents used/water also added/added in addition or smaller volumes therefore greater percentage error	1
		4
	Total:	11

Page 7	Mark Scheme	Syllabus	Paper
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Question	Answer	Marks																									
FA 7 is ZnSO ₄ ; FA 8 is (NH ₄) ₂ Fe(SO ₄) ₂ ; FA 9 is CrK(SO ₄) ₂ ; FA 10 is MnSO ₄ ; FA 11 is NaNO ₂																											
3(a)	<p>Selects NaOH and NH₃</p> <p>Single table to show results with both NaOH and NH₃. No repeat headings. At least two of the FAs tested</p> <table border="1" style="margin-left: auto; margin-right: auto;"> <thead> <tr> <th></th> <th>FA 7</th> <th>FA 8</th> <th>FA 9</th> <th>FA 10</th> </tr> </thead> <tbody> <tr> <td>NaOH</td> <td>white ppt</td> <td>green ppt</td> <td>grey-green ppt</td> <td>off-white/ pale brown/ buff ppt</td> </tr> <tr> <td>excess</td> <td>soluble</td> <td>insoluble</td> <td>soluble</td> <td>insoluble</td> </tr> <tr> <td>NH₃</td> <td>white ppt</td> <td>green ppt</td> <td>grey-green ppt</td> <td>off-white/ pale brown/ buff ppt</td> </tr> <tr> <td>excess</td> <td>soluble</td> <td>insoluble</td> <td>insoluble</td> <td>insoluble</td> </tr> </tbody> </table> <p>FA 8 ppt turning brown in air with either NaOH or NH₃</p> <p>FA 9 ppt dissolves to form (dark) green solution with excess NaOH</p> <p>FA 10 darkens/goes brown with either NaOH or NH₃</p>		FA 7	FA 8	FA 9	FA 10	NaOH	white ppt	green ppt	grey-green ppt	off-white/ pale brown/ buff ppt	excess	soluble	insoluble	soluble	insoluble	NH ₃	white ppt	green ppt	grey-green ppt	off-white/ pale brown/ buff ppt	excess	soluble	insoluble	insoluble	insoluble	<p>1</p> <p>1</p> <p>1</p> <p>1</p> <p>1</p> <p>1</p> <p>1</p> <p>1</p> <p>1</p> <p>1</p>
	FA 7	FA 8	FA 9	FA 10																							
NaOH	white ppt	green ppt	grey-green ppt	off-white/ pale brown/ buff ppt																							
excess	soluble	insoluble	soluble	insoluble																							
NH ₃	white ppt	green ppt	grey-green ppt	off-white/ pale brown/ buff ppt																							
excess	soluble	insoluble	insoluble	insoluble																							

Page 8	Mark Scheme	Syllabus	Paper
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Question	Answer	Marks								
	<table border="1"> <tr> <td>FA 7</td> <td>FA 8</td> <td>FA 9</td> <td>FA 10</td> </tr> <tr> <td>Zn²⁺</td> <td>Fe²⁺</td> <td>Cr³⁺</td> <td>Mn²⁺</td> </tr> </table> <p>Award 1 mark for 2 correct ions. Award 2 marks for all 4 correct.</p>	FA 7	FA 8	FA 9	FA 10	Zn ²⁺	Fe ²⁺	Cr ³⁺	Mn ²⁺	<p>1 1</p> <p>11</p>
FA 7	FA 8	FA 9	FA 10							
Zn ²⁺	Fe ²⁺	Cr ³⁺	Mn ²⁺							
3(b)	<p>(dark) brown ppt/solid/suspension/deposit and effervescence/bubbling/fizzing</p> <p>positive test for oxygen – (gas/ O₂) relights glowing splint</p>	<p>1</p> <p>1</p> <p>2</p>								
3(c)(i)	<p>blue solution and effervescence/bubbling/fizzing or brown fumes/gas</p>	1								
3(c)(ii)	<p>NO₂⁻ or nitrite from either blue solution or brown gas</p>	1								

Page 9	Mark Scheme	Syllabus	Paper
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Question	Answer	Marks
3(c)(iii)	<p>selects NaOH and Al (for nitrite or nitrate) or selects (acidified) potassium manganate(VII)/ potassium permanganate/ KMnO_4</p> <p>If carbonate in (ii) (from bubbling without brown gas in (i)) then allow use of limewater to test gas If halide from no reaction then allow use of AgNO_3 and NH_3 If sulfate/sulfite from no reaction then allow use of $\text{BaCl}_2/\text{Ba}(\text{NO}_3)_2$ and HCl/HNO_3</p> <p>Warming (with NaOH and Al) and gas/ammonia turns (damp red) litmus (paper) blue or Decolourises MnO_4^-</p>	<p>1</p> <p>1</p> <p>4</p>
	Total:	17